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(a) Sample answer:  $q = m \times C \times \Delta. T$   $q = 210 \times 4.18$   
 $\times 65$  .  $q = 57057$  .  $q = 57.057$  kJ . Now  $\Delta. H. Comb =$   
 $q n$  .  $57.057$  .  $n. C = 2. H. 5. OH$  .  $1367$  .  $n = 0.04174$   
mol. C. 2. H. 5. OH  $\therefore m. C = n \times MM. 2. H. 5. OH$  C. 2.

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H. 5. OH  $\therefore$  m. C =  $0.04174 \times 46.068$ . 2. H. 5. OH  $\therefore$  m.  
C. 2. H. 5. OH = 1.923 g . Now 50% of heat has been  
lost to surroundings.  $\therefore$  initial . m =  $2 \times 1.920$  g . C. 2.  
H. 5

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