

Physical Chemistry Thermodynamics Problems And Solutions

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Thermodynamics Problem Solving in Physical Chemistry ...
Correct answer: I and III. Explanation: Condition I is always true. Condition II is never true, as Gibbs free energy cannot be negative if enthalpy is positive and entropy is negative. Condition III may be true if temperature is very high (this is the scenario when the term dominates the term.

Thermochemistry and Thermodynamics - Physical Chemistry
Problem : Given that the free energy of formation of liquid water is -237 kJ / mol, calculate the potential for the formation of hydrogen and oxygen from water. To solve this problem we must first calculate ΔG for the reaction, which is -2 (-237 kJ / mol) = 474 kJ / mol. Knowing that $\Delta G = -nFE$ and $n = 4$, we calculate the potential is -1.23 V.

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Text: Atkins and de Paula, Physical Chemistry 9th edition (used copy/online bookseller) Website: learn@UW for notices, problem sets, exams, handouts, power points, answer keys Tentative Course Outline Thermodynamics is a " theory of everything " , and is one of the most beautiful branches of science we know, touching every aspect of our lives.

Physical Chemistry: Thermodynamics and Kinetics ...
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